

ANEXOS

**ANEXO A: ESPECIFICACION DE EQUIPOS
Y MATERIALES UTILIZADOS EN LA
EXPERIMENTACIÓN**

BALANZA ELÉCTRICA



Modelo	EU-C500
Industria	Italia
Capacidad	510 g
Rango de Temperatura	10-40 °C
Error de Precisión	0,001 g
Tensión	120 V
Frecuencia	50 Hz

ESTUFA DE SECADO



Modelo	Conterm 2000201
Marca	J.P Selecta S.A.
Rango de Temperatura	40-250 °C
Potencia	2000 w
Voltaje	230 v
Frecuencia	50/60 Hz

MUFLA



Modelo	VEN MKH
Marca	VEB elektro bad frankenhausen
Rango de Temperatura	10-1200 °C
Potencia	3,2 kw
Voltaje	220 v
Frecuencia	50 Hz

ESPECTROFOTOMETRO



Modelo	VR-2000
Marca	J.P. Selecta, s.a
Rango de onda	320-1000 nm
Corriente	112mA
Voltaje	230/240 v
Frecuencia	50-60 Hz

SECADOR DE INFRARROJOS



Modelo	Eurotherm
Marca	Sartorius
Potencia	350 VA
Rango de temperatura	10-40 °C
Voltaje	220 v
Frecuencia	50 Hz

TAMIZADOR



Marca	ORTO-ALRESA
Potencia	100 W
Velocidad	2500 rpm
Voltaje	220 v
Frecuencia	50 - 60 Hz
Recorrido de tamices	3mm- 4um

**ANEXO B: METODO ESTANDAR DE
DETERMINACION DE ÍNDICE DE YODO
DEL CARBÓN ACTIVADO (ASTM D-4607)**



Designation: D 4607 – 94 (Reapproved 2006)

Standard Test Method for Determination of Iodine Number of Activated Carbon¹

This standard is issued under the fixed designation D 4607; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This test method covers the determination of the relative activation level of unused or reactivated carbons by adsorption of iodine from aqueous solution. The amount of iodine absorbed (in milligrams) by 1 g of carbon using test conditions listed herein is called the iodine number.

1.2 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use. Specific hazard statements are given in Section 7.*

2. Referenced Documents

2.1 *ASTM Standards:*²

C 819 Test Method for Specific Surface Area of Carbon or Graphite

D 1193 Specification for Reagent Water

D 2652 Terminology Relating to Activated Carbon

D 2867 Test Methods for Moisture in Activated Carbon

D 3860 Practice for Determination of Adsorptive Capacity of Activated Carbon by Aqueous Phase Isotherm Technique

E 11 Specification for Wire Cloth and Sieves for Testing Purposes

E 177 Practice for Use of the Terms Precision and Bias in ASTM Test Methods

E 287 Specification for Laboratory Glass Graduated Burets

E 288 Specification for Laboratory Glass Volumetric Flasks

E 300 Practice for Sampling Industrial Chemicals

2.2 *NIST Publication:*

¹ This test method is under the jurisdiction of ASTM Committee D28 on Activated Carbon and is the direct responsibility of Subcommittee D28.02 on Liquid Phase Evaluation.

Current edition approved Oct. 1, 2006. Published November 2006. Originally approved in 1986. Last previous edition approved in 1999 as D 4607 – 94 (1999).

² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For Annual Book of ASTM Standards volume information, refer to the standard's Document Summary page on the ASTM website.

Circular 602—Testing of Glass Volumetric Apparatus³

3. Summary of Test Method

3.1 This test method is based upon a three-point adsorption isotherm (see Practices D 3860). A standard iodine solution is treated with three different weights of activated carbon under specified conditions. The carbon treated solutions are filtered to separate the carbon from the treated iodine solution (filtrate). Iodine remaining in the filtrate is measured by titration. The amount of iodine removed per gram of carbon is determined for each carbon dosage and the resulting data used to plot an adsorption isotherm. The amount of iodine adsorbed (in milligrams) per gram of carbon at a residual iodine concentration of 0.02 *N* is reported as the iodine number.

3.2 Iodine concentration in the standard solution affects the capacity of an activated carbon for iodine adsorption. Therefore, the normality of the standard iodine solution must be maintained at a constant value (0.100 ± 0.001 *N*) for all iodine number measurements.

3.3 The apparatus required consists of various laboratory glassware used to prepare solutions and contact carbon with the standard iodine solution. Filtration and titration equipment are also required.

4. Significance and Use

4.1 The iodine number is a relative indicator of porosity in an activated carbon. It does not necessarily provide a measure of the carbon's ability to absorb other species. Iodine number may be used as an approximation of surface area for some types of activated carbons (see Test Method C 819). However, it must be realized that any relationship between surface area and iodine number cannot be generalized. It varies with changes in carbon raw material, processing conditions, and pore volume distribution (see Definitions D 2652).

4.2 The presence of adsorbed volatiles, sulfur, and water extractables may affect the measured iodine number of an activated carbon.

³ Available from National Institute of Standards and Technology (NIST), 100 Bureau Dr., Stop 1070, Gaithersburg, MD 20899-1070, <http://www.nist.gov>.

5. Apparatus

Note 1—All volumetric measuring equipment should meet or exceed the requirements of NIST Circular 602. Volumetric glassware meeting these specifications is generally designated as "Class A". See also Specifications E 287 and E 288.

- 5.1 Analytical Balance, accuracy ± 0.0001 g.
- 5.2 Buret, 10-mL capacity or 5-mL precision buret.
- 5.3 Flasks, Erlenmeyer 250-mL capacity with ground glass stoppers.
- 5.4 Flask, Erlenmeyer wide-mouthed, 250-mL capacity.
- 5.5 Beakers, assorted sizes.
- 5.6 Bottles, amber, for storage of iodine and thiosulfate solutions.
- 5.7 Funnel, 100-mm top inside diameter.
- 5.8 Filter Paper, 18.5-cm prefolded paper, Whatman No. 2V or equivalent.
- 5.9 Pipets, volumetric type, 5.0, 10.0, 25.0, 50.0, and 100.0-mL capacity.
- 5.10 Volumetric Flasks, 1 L.
- 5.11 Graduated Cylinders, 100 mL and 500 mL.

6. Reagents

6.1 **Purity of Reagents**—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.⁴ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

6.2 **Purity of Water**—References to water shall be understood to mean reagent water conforming to Specification D 1193 for Type II reagent water.

- 6.3 Hydrochloric Acid, concentrated.
- 6.4 Sodium Thiosulfate, $(\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O})$.
- 6.5 Iodine, United States Pharmacopeia, resublimed crystals.
- 6.6 Potassium Iodide.
- 6.7 Potassium Iodate, primary standard.
- 6.8 Starch, soluble potato or arrowroot.
- 6.9 Sodium Carbonate.

7. Hazards

7.1 Several potential hazards are associated with conducting this test procedure. It is not the purpose of this standard to address all potential health and safety hazards encountered with its use. The user is responsible for establishing appropriate health and safety practices before use of this test procedure. Determine the applicability of federal and state regulations before attempting to use this test method.

7.2 Personnel conducting the iodine number procedure should be aware of potential safety and health hazards associ-

ated with the chemicals used in this procedure. The "Material Safety Data Sheet" (MSDS) for each reagent listed in Section 6 should be read and understood. Special precautions to be taken during use of each reagent are included on the "Material Safety Data Sheet" (MSDS). First aid procedures for contact with a chemical are also listed on its "MSDS." A "Material Safety Data Sheet" for each reagent may be obtained from the manufacturer. Other safety and health hazard information on reagents used in this procedure is available.^{5,6,7}

7.3 Careful handling and good laboratory technique should always be used when working with chemicals. Avoid contact with hydrochloric acid or acid vapor. Care should also be taken to prevent burns during heating of various solutions during this test procedure.

7.4 The user of this test method should comply with federal, state, and local regulations for safe disposal of all samples and reagents used.

8. Preparation of Solutions

8.1 **Hydrochloric Acid Solution (5% by weight)**—Add 70 mL of concentrated hydrochloric acid to 550 mL of distilled water and mix well. A graduated cylinder may be used for measurement of volume.

8.2 **Sodium Thiosulfate (0.100 N)**—Dissolve 24.820 g of sodium thiosulfate in approximately 75 \pm 25 mL of freshly boiled distilled water. Add 0.10 \pm 0.01 g of sodium carbonate to minimize bacterial decomposition of the thiosulfate solution. Quantitatively transfer the mixture to a 1-L volumetric flask and dilute to the mark. Allow the solution to stand at least 4 days before standardizing. The solution should be stored in an amber bottle.

8.3 **Standard Iodine Solution (0.100 \pm 0.001 N)**—Weigh 12.700 g of iodine and 19.100 g of potassium iodide (KI) into a beaker. Mix the dry iodine and potassium iodide. Add 2 to 5 mL of water to the beaker and stir well. Continue adding small increments of water (approximately 5 mL each) while stirring until the total volume is 50 to 60 mL. Allow the solution to stand a minimum of 4 h to ensure that all crystals are thoroughly dissolved. Occasional stirring during this 4-h period will aid in the dissolution. Quantitatively transfer to a 1-L volumetric flask and fill to the mark with distilled water. It is important that the standard iodine solution has an iodide-to-iodine weight ratio of 1.5 to 1. Store the solution in an amber bottle.

8.4 **Potassium Iodate Solution (0.1000 N)**—Dry 4 or more grams of primary standard grade potassium iodate (KIO_3) at $110 \pm 5^\circ\text{C}$ for 2 h and cool to room temperature in a desiccator. Dissolve 3.5667 \pm 0.1 mg of the dry potassium iodate in about 100 mL of distilled water. Quantitatively transfer to a 1-L volumetric flask and fill to the mark with distilled water. Mix thoroughly and store in a glass-stoppered bottle.

⁴ Reagent Chemicals, American Chemical Society Specifications, "Am. Chem. Soc., Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see "Reagent Chemicals and Standards," by Joseph Rosin, D. Van Nostrand Co., Inc., New York, NY, and the "United States Pharmacopeia."

⁵ The "Chemical Safety Data Sheet" for the subject chemical is available from the Manufacturing Chemists Association, Washington, DC.

⁶ Sax, N. I., *Dangerous Properties of Industrial Materials*, 4th edition, 1975, Van Nostrand Reinhold Company, New York, NY.

⁷ *NIOSH/OSHA Pocket Guide to Chemical Hazards*, 1978, U.S. Department of Labor, Occupational Safety and Health Administration, Washington, DC. Available from U.S. Government Printing Office, Washington, DC.

8.5 *Starch Solution*—Mix 1.0 ± 0.5 g of starch with 5 to 10 mL of cold water to make a paste. Add an additional 25 ± 5 mL of water while stirring to the starch paste. Pour the mixture, while stirring, into 1 L of boiling water and boil for 4 to 5 min. This solution should be made fresh daily.

9. Standardization of Solutions

9.1 *Standardization of 0.100 N Sodium Thiosulfate*—Pipet 25.0 mL of potassium iodate (KIO_3) solution from 8.4 into a 250-mL titration (or wide-mouthed Erlenmeyer) flask. Add 2.00 ± 0.01 g of potassium iodide (KI) to the flask and shake the flask to dissolve the potassium iodide crystals. Pipet 5.0 mL of concentrated hydrochloric acid into the flask. Titrate the free iodine with sodium thiosulfate solution until a light yellow color is observed in the flask. Add a few drops of starch indicator (8.5) and continue the titration dropwise until one drop produces a colorless solution. Determine sodium thiosulfate normality as follows:

$$N_1 = (P/R)S \quad (1)$$

where:

N_1 = sodium thiosulfate, N ,
 P = potassium iodate, mL,
 R = potassium iodate, N , and
 S = sodium thiosulfate, mL.

The titration step should be done in triplicate and the normality results averaged. Additional replications should be done if the range of values exceeds 0.003 N .

9.2 *Standardization of 0.100 \pm 0.001 N Iodine Solution*—Pipet 25.0 mL of iodine solution (8.3) into a 250-mL wide-mouthed Erlenmeyer flask. Titrate with standardized sodium thiosulfate (9.1) until the iodine solution is a light yellow color. Add a few drops of starch indicator and continue titration dropwise until one drop produces a colorless solution. Determine the iodine solution normality as follows:

$$N_2 = (S - N_1)I \quad (2)$$

where:

N_2 = iodine, N ,
 S = sodium thiosulfate, mL,
 N_1 = sodium thiosulfate, N , and
 I = iodine, mL.

The titration step should be done in triplicate and the normality results averaged. Additional replications should be done if the range of values exceeds 0.003 N . The iodine solution concentration must be $0.100 \pm 0.001 N$. If this requirement is not met, repeat 8.3 and 9.2.

10. Procedure

10.1 The procedure applies to either powdered or granular activated carbon. When granular carbon is to be tested, grind a representative sample (see Practice E 300) of carbon until 60 wt % (or more) will pass through a 325-mesh screen and 95 wt % (or more) will pass through a 100-mesh screen (U.S. sieve series, see Specification E 11). Carbon received in the powdered form may need additional grinding to meet the particle size requirement given above.

10.2 Dry the ground carbon from 10.1 in accordance with Test Method D 2867. Cool the dry carbon to room temperature in a desiccator.

10.3 Determination of iodine number requires an estimation of three carbon dosages. Section 11.4 describes how to estimate the carbon dosages to be used. After estimating carbon dosages, weigh three appropriate amounts of dry carbon to the nearest milligram. Transfer each weighed sample of carbon to a clean, dry 250-mL Erlenmeyer flask equipped with a ground glass stopper.

10.4 Pipet 10.0 mL of 5 wt % hydrochloric acid solution into each flask containing carbon. Stopper each flask and swirl gently until the carbon is completely wetted. Loosen the stoppers to vent the flasks, place on a hot plate in a fume hood, and bring the contents to a boil. Allow to boil gently for 30 ± 2 s to remove any sulfur which may interfere with the test results. Remove the flasks from the hot plate and cool to room temperature.

10.5 Pipet 100.0 mL of 0.100 N iodine solution into each flask. Standardize the iodine solution just prior to use. Stagger the addition of iodine to the three flasks so that no delays are encountered in handling. Immediately stopper the flasks, and shake the contents vigorously for 30 ± 1 s. Quickly filter each mixture by gravity through one sheet of folded filter paper (Whatman No. 2V or equivalent) into a beaker. Filtration equipment must be prepared in advance so no delay is encountered in filtering the samples.

10.6 For each filtrate, use the first 20 to 30 mL to rinse a pipet. Discard the rinse portions. Use clean beakers to collect the remaining filtrates. Mix each filtrate by swirling the beaker and pipet 50.0 mL of each filtrate into a clean 250-mL Erlenmeyer flask. Titrate each filtrate with standardized 0.100 N sodium thiosulfate solution until the solution is a pale yellow. Add 2 mL of the starch indicator solution and continue the titration with sodium thiosulfate until one drop produces a colorless solution. Record the volume of sodium thiosulfate used.

11. Calculation

11.1 The capacity of a carbon for any adsorbate is dependent upon the concentration of the adsorbate in solution. The concentrations of the standard iodine solution and filtrates must be specified or known. This is necessary to determine an appropriate carbon weight to produce final concentrations agreeing with the definition of iodine number. The amount of carbon sample to be used in the determination is governed by the activity of the carbon. If filtrate normalities (C) are not within the range of 0.008 N to 0.040 N , repeat the procedure using different carbon weights.

11.2 Two calculations are required for each carbon dosage, as X/M and C .

11.2.1 To calculate the value of X/M , first derive the following values:

$$A = (N_2)(12693.0) \quad (3)$$

where:

N_2 = iodine, N (from 9.2).

$$B = (N_s)(126.93) \quad (4)$$

where:

N_s = sodium thiosulfate, N (from 9.1).

$$DF = (I + H)/F \quad (5)$$

where:

DF = dilution factor,

I = iodine, mL (from 9.2),

H = 5 % hydrochloric acid used, mL, and

F = filtrate, mL.

For example, if 10 mL of HCl and 50 mL of filtrate are used:

$$DF = (10 + 10)/50 = 2.2$$

11.2.1.1 Calculate the value of X/M as follows:

$$X/M = [A - (DF)(B)(S)]/M \quad (6)$$

where:

X/M = iodine absorbed per gram of carbon, mg/g,

S = sodium thiosulfate, mL, and

M = carbon used, g.

11.2.2 Calculate the value of C as follows:

$$C = (N_r S)/F \quad (7)$$

where:

C = residual filtrate, N ,

N_r = sodium thiosulfate, N , and

F = filtrate, mL.

11.3 Using logarithmic paper, plot X/M (as the ordinate) versus C (as the abscissa) for each of the three carbon dosages (see Fig. 1). Calculate the least squares fit for the three points and plot. The iodine number is the X/M value at a residual iodine concentration (C) of 0.02 N . The regression coefficient for the least squares fit should be greater than 0.995.

11.4 Carbon dosage may be estimated as follows:

$$M = [A - (DF)(C)(126.93)(S)]/E \quad (8)$$

where:

M = carbon, g,

A = $(N_s)(12693.0)$,

DF = dilution factor (see 11.2.1),

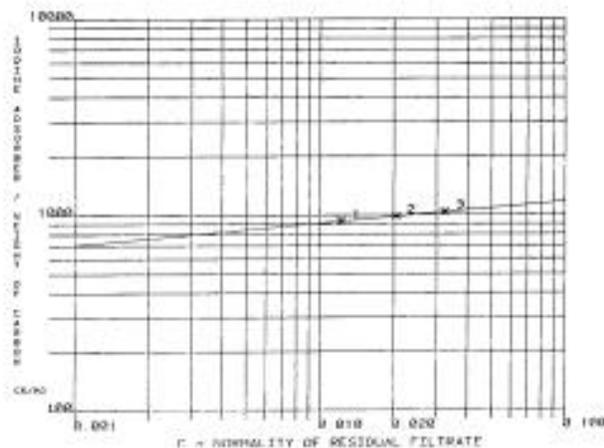
C = residual iodine, and

E = estimated iodine number of the carbon.

Three carbon dosages are calculated using three values of C (usually 0.01, 0.02, and 0.03).

12. Report

12.1 The reports should include the following:



LAB NO. 4 OCT '80-3
TDS NO. 804 PAGE 22
TECH P/SP

POINT 1		
X/M	--	932
C	--	0.012
POINT 2		
X/M	--	964
C	--	0.020
POINT 3		
X/M	--	1040
C	--	0.030
IODINE NO.	--	964
SLOPE	--	0.111
CORR. COEF	--	0.999

FIG. 1 Activated Carbon Iodine Adsorption Isotherm

2.3 IODINE ADSORPTION

2.3.1. Scope

The determination of the iodine number is a simple and quick test, giving an indication of the internal surface area of activated carbon.

2.3.2. Principle

The iodine number is defined as the number of milligrams of iodine adsorbed from an aqueous solution by 1 g of activated carbon when the iodine concentration of the residual filtrate is 0.02 N.

2.3.3. Apparatus and materials

Glass stoppered flasks, pipettes, burette,

Folded filters, Whatman 2V, 18 cm Ø,

Hot-plate,

Hydrochloric acid, 5 per cent (w/w),

Sodium thiosulphate solution, 0.10 N, standardized,

Iodine solution, 0.10 N, standardized,

Starch solution.

2.3.4. Sample

Granular activated carbon is pulverized (< 0.1 mm) and then dried at 150 °C to constant weight.

2.3.5. Procedure

Depending on the activity of the carbon, weigh 0.700 — 2.000 g of the dried carbon and transfer the weighed sample to a dry, glass stoppered 250 ml Erlenmeyer flask.

Pipette 10 ml of 5% HCl into the flask and swirl until the activated carbon is wetted.

Place the flask on a hot-plate, bring the contents to the boil, and allow to boil for exactly 30 seconds.

Allow the flask and contents to cool to room temperature, then add by pipette 100 ml of 0.10 N iodine solution.

Stopper the flask immediately and shake it vigorously for 30 seconds. Filter by gravity through a filter paper immediately after the 30 seconds shaking period.

Discard the initial 20 — 30 ml of filtrate and collect the remainder in a clean beaker.

Stir the filtrate in the beaker with a glass rod and pipette 50 ml into a 250 ml Erlenmeyer flask.

Titrate the 50 ml sample with 0.10 N sodium thiosulphate solution until the yellow colour has almost disappeared. Add about 1 ml of starch solution and continue titration until the blue indicator colour just disappears. Record the volume of sodium thiosulphate solution used.

2.3.6. Result

Calculate the iodine number I_n of the carbon using the equation:

$$I_n = \frac{X}{M} A$$

where

X = mg of iodine adsorbed by the activated carbon

$$X = (12.693 N_1) - (279.246 N_2 V)$$

herein

N_1 = normality of iodine solution

N_2 = normality of sodium thiosulphate solution

V = volume of sodium thiosulphate solution in ml

M = mass of activated carbon in g

A = correction factor, depending on the residual normality N_r of the filtrate

This factor A may be applied if N_r is between 0.008 and 0.0334 N.

$$N_r = N_2 \frac{V}{50}$$

A is given in the table below. A may vary from 1.1625 to 0.925. If N_r is outside the range of 0.008 to 0.0334 N, the determination has to be repeated either with a larger amount of activated carbon if $N_r > 0.0334$ or a smaller amount if $N_r < 0.008$.

Reference:

AWWA B 600-78 Powdered Activated Carbon.

Extracts were reprinted from B600-78- *AWWA Standard for Activated Carbon*, by permission.

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**ANEXO C: TABLA DE FACTOR DE
CORRECCIÓN PARA N° DE YODO**

NORMALIDAD RESIDUAL FILTRADO (C)	0.0000	0.0001	0.0002	0.0003	0.0004	0.0005	0.0006	0.0007	0.0008	0.0009
0.0080	1,1625	1,1613	1,1600	1,1575	1,1550	1,1538	1,1530	1,1500	1,1475	1,1463
0.0090	1,1438	1,1425	1,1400	1,1375	1,1363	1,1350	1,1325	1,1300	1,1288	1,1275
0.0100	1,1250	1,1238	1,1225	1,1213	1,1200	1,1175	1,1163	1,1150	1,1138	1,1113
0.0110	1,1100	1,1088	1,1075	1,1063	1,1038	1,1025	1,1000	1,0988	1,0975	1,0963
0.0120	1,0950	1,0938	1,0925	1,0900	1,0888	1,0875	1,0863	1,0850	1,0838	1,0825
0.0130	1,0800	1,0788	1,0775	1,0763	1,0750	1,0738	1,0725	1,0713	1,0700	1,0688
0.0140	1,0675	1,0663	1,0650	1,0625	1,0613	1,0600	1,0588	1,0575	1,0563	1,0550
0.0150	1,0538	1,0525	1,0513	1,0500	1,0488	1,0475	1,0453	1,045	1,0425	1,0425
0.0160	1,0413	1,0400	1,0388	1,0375	1,0375	1,0363	1,0350	1,0333	1,0325	1,0313
0.0170	1,0300	1,0288	1,0275	1,0263	1,0250	1,0245	1,0238	1,0225	1,0203	1,0200
0.0180	1,0200	1,0188	1,0175	1,0163	1,0150	1,0144	1,0138	1,0125	1,0125	1,0113
0.0190	1,0100	1,0088	1,0075	1,0075	1,0063	1,0050	1,0050	1,0038	1,0025	1,0025
0.0200	1,0013	1,0000	1,0000	0,9988	0,9750	0,9975	0,9963	0,9950	0,9950	0,9938
0.0210	0,9938	0,9925	0,9925	0,9913	0,9900	0,9900	0,9888	0,9875	0,9875	0,9963
0.0220	0,9863	0,9850	0,9850	0,9838	0,9825	0,9825	0,9813	0,9813	0,9800	0,9788
0.0230	0,9788	0,9775	0,9775	0,9763	0,9763	0,9750	0,9750	0,9738	0,9738	0,9725
0.0240	0,9725	0,9708	0,9700	0,9700	0,9688	0,9688	0,9675	0,9675	0,9663	0,9663
0.0250	0,9650	0,9650	0,9638	0,9638	0,9625	0,9625	0,9613	0,9613	0,9606	0,9600
0.0260	0,9600	0,9588	0,9588	0,9575	0,957	0,9563	0,9563	0,9550	0,9550	0,9538
0.0270	0,9538	0,9525	0,9525	0,9519	0,9513	0,9513	0,9506	0,9500	0,9500	0,9438
0.0280	0,9488	0,9475	0,9475	0,9463	0,9463	0,9463	0,9450	0,9450	0,9438	0,9438
0.0290	0,9425	0,9425	0,9425	0,9413	0,9413	0,9400	0,9400	0,9394	0,9388	0,9388
0.0300	0,9375	0,9375	0,9375	0,9363	0,9363	0,9363	0,9363	0,9350	0,9350	0,9346
0.0310	0,9333	0,9333	0,9325	0,9325	0,9325	0,9319	0,9319	0,9313	0,9300	0,9300
0.0320	0,9300	0,9294	0,9288	0,9288	0,9280	0,9275	0,9275	0,9275	0,9270	0,9270
0.0330	0,9263	0,9263	0,9257	0,9250	0,9250					

**ANEXO D: DETERMINACIÓN DEL ÍNDICE
DE AZUL DE METILENO**

2.4. METHYLENE BLUE VALUE

2.4.1. Scope

The methylene blue value gives an indication of the adsorption capacity of an activated carbon for molecules having similar dimensions to methylene blue. It is a quick method for comparing different batches of activated carbon of the same quality.

2.4.2. Principle

The methylene blue value is defined as the number of millilitres standard methylene blue solution decolourized by 0.1 g of activated carbon (dry basis).

2.4.3. Apparatus and materials

Glass stoppered flask.

Filters.

Methylene blue test solution :

Dissolve a quantity, equivalent to 1200 mg of pure dye (approx. 1.5 g Methyleneblau DAB VI or equivalent *) to 1000 ml in a volumetric flask. Allow the solution to stand several hours or overnight.

Check the solution by diluting 5.0 ml with 0.25% (v/v) acetic acid to 1 l in a volumetric flask and measuring the absorbance at 620 nm for 1 cm. The absorbance must be 0.840 ± 0.01 .

If the absorbance is higher, dilute with the calculated amount of water. If lower, discard the solution and start over.

2.4.4. Sample

Granular activated carbon is pulverized (< 0.1 mm) and then dried at 150 °C to constant weight.

2.4.5. Procedure

Contact exactly 0.1 g of the carbon sample with 25 (5)** ml of the methylene blue test solution in a glass stoppered flask. Shake until decolourization occurs. Then add a further 5 (1)** ml of the methylene blue test solution and shake to decolourization. Repeat the addition of methylene blue test solution in 5 (1) ml portions as long as decolourization occurs within five minutes.

Note the entire volume of test solution decolourized by the sample.

Repeat the test to confirm the result obtained.

ANEXO E: ESTIMACIÓN DE COSTOS

EVALUACIÓN DE COSTOS

Los costos que conlleva la realización del presente proyecto de investigación se muestran a continuación en las siguientes tablas:

Costo de Materia Prima

Detalle	Costo (Bs.)
Borras de café (1kg)	0

Costo de reactivos utilizados en la parte experimental y análisis

N°	Reactivos	Detalle	Costo (Bs.)
1	Yodo 0,1 N	2 litros	352
2	Tiosulfato de sodio 0,1 N	½ litro	120
3	Ácido fosfórico al 85%	½ litro	30
4	Ácido clorhídrico al 5%	200 ml	34
5	Hidróxido de sodio	½ litro	90
6	Azul de metileno al 1%	½ litro	224
7	Almidón	250ml	25
8	Agua destilada	6 litros	48
Total			923

Costo de materiales adquiridos para realizar la parte experimental

N°	Detalle	Descripción	Costo (Bs.)
1	Crisol de 100 ml	1 unidad	42
2	Papel filtro Whatman 42	3 pliegues	15
3	Frascos de 50 ml	16 unidades	24
4	Botellas de 500 ml	5 unidades	15
Total			96

Costo de materiales directos e indirectos

N°	Detalle	Descripción	Costo (Bs.)
1	Impresión	3 ejemplares	210
2	Empastados	3 ejemplares	180
Total			390

Costos energéticos

N°	Detalle	Energía consumida (Kwh)	Costo de Kwh (Bs.)	Costo (Bs.)
1	Secado	28	0,90	25,2
2	Pre carbonización	6,4	0,90	5,76
3	Carbonización	16	0,90	14,4
Total				45,36

Costo total

Costo Total	Costo (Bs.)
Costo de materia prima	0
Costo de reactivos utilizados en la parte experimental y análisis	923
Costo de materiales adquiridos para realizar la parte experimental	96
Costo de materiales directos e indirectos	390
Costos energéticos	45,36
Total	1454,36